

Location Entry Codes



As part of CIE's continual commitment to maintaining best practice in assessment, CIE uses different variants of some question papers for our most popular assessments with large and widespread candidature. The question papers are closely related and the relationships between them have been thoroughly established using our assessment expertise. All versions of the paper give assessment of equal standard.

The content assessed by the examination papers and the type of questions is unchanged.

This change means that for this component there are now two variant Question Papers, Mark Schemes and Principal Examiner's Reports where previously there was only one. For any individual country, it is intended that only one variant is used. This document contains both variants which will give all Centres access to even more past examination material than is usually the case.

The diagram shows the relationship between the Question Papers, Mark Schemes and Principal Examiners' Reports that are available.

Question Paper	Mark Scheme	Principal Examiner's Report
Introduction	Introduction	Introduction
First variant Question Paper	First variant Mark Scheme	First variant Principal Examiner's Report
Second variant Question Paper	Second variant Mark Scheme	Second variant Principal Examiner's Report

Who can I contact for further information on these changes?

Please direct any questions about this to CIE's Customer Services team at:

international@cie.org.uk

The titles for the variant items should correspond with the table above, so that at the top of the first page of the relevant part of the document and on the header, it has the words:

- First variant Question Paper / Mark Scheme / Principal Examiner's Report

or

- Second variant Question Paper / Mark Scheme / Principal Examiner's Report

as appropriate.



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CANDIDATE
NAME

CENTRE
NUMBER

--	--	--	--	--

CANDIDATE
NUMBER

--	--	--	--

* 2 9 4 2 3 9 7 2 8 8 *

CHEMISTRY

0620/31

Paper 3 (Extended)

October/November 2008

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES

Answer **all** questions.

A copy of the Periodic Table is printed on page 12.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part questions.

For Examiner's Use	
1	
2	
3	
4	
5	
6	
7	
Total	

This document consists of **12** printed pages.



1 Complete the following table.

gas	test for gas
ammonia	
	bleaches damp litmus paper
hydrogen	
	relights a glowing splint
	turns limewater milky

[Total: 5]

*For
Examiner's
Use*

2 There are three types of giant structure – ionic, metallic and macromolecular.

(a) Sodium nitride is an ionic compound. Draw a diagram that shows the formula of the compound, the charges on the ions and the arrangement of the valency electrons around the negative ion.

Use x to represent an electron from a sodium atom.
Use o to represent an electron from a nitrogen atom.

[3]

(b) (i) Describe metallic bonding.

.....
..... [3]

(ii) Use the above ideas to explain why
metals are good conductors of electricity,

..... [1]

metals are malleable.

..... [2]

(c) Silicon(IV) oxide has a macromolecular structure.

(i) **Describe** the structure of silicon(IV) oxide (a diagram is not acceptable).

.....
.....
..... [3]

(ii) Diamond has a similar structure and consequently similar properties.
Give **two** physical properties common to both diamond and silicon(IV) oxide.

.....
..... [2]

[Total: 14]

3 Steel is an alloy made from impure iron.

(a) Both iron and steel rust. The formula for rust is $\text{Fe}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$.
It is hydrated iron(III) oxide.

(i) Name the **two** substances that must be present for rusting to occur.

..... [2]

(ii) Painting and coating with grease are two methods of preventing iron or steel from rusting. Give **two** other methods.

.....
..... [2]

(b) (i) Name a reagent that can reduce iron(III) oxide to iron.

..... [1]

(ii) Write a symbol equation for the reduction of iron(III) oxide, Fe_2O_3 , to iron.

..... [2]

(c) (i) Calculate the mass of one mole of $\text{Fe}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$.

..... [1]

(ii) Use your answer to (i) to calculate the percentage of iron in rust.

.....
..... [2]

(d) Iron from the blast furnace is impure. Two of the impurities are carbon and silicon. These are removed by blowing oxygen through the molten iron and adding calcium oxide.

(i) Explain how the addition of oxygen removes carbon.

.....
..... [1]

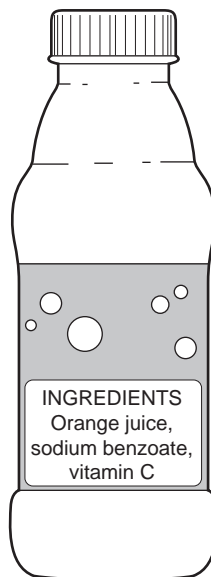
(ii) Explain how the addition of oxygen and calcium oxide removes silicon.

.....
..... [2]

[Total: 13]

- 4 Across the world, food safety agencies are investigating the presence of minute traces of the toxic hydrocarbon, benzene, in soft drinks. It is formed by the reduction of sodium benzoate by vitamin C.

For
Examiner's
Use



- (a) Sodium benzoate is a salt, it has the formula C_6H_5COONa . It can be made by the neutralisation of benzoic acid by sodium hydroxide.

(i) Deduce the formula of benzoic acid.

..... [1]

(ii) Write a word equation for the reaction between benzoic acid and sodium hydroxide.

..... [1]

(iii) Name **two** other compounds that would react with benzoic acid to form sodium benzoate.

..... [2]

- (b) Benzene contains 92.3% of carbon and its relative molecular mass is 78.

(i) What is the percentage of hydrogen in benzene?

..... [1]

(ii) Calculate the ratio of moles of C atoms: moles of H atoms in benzene.

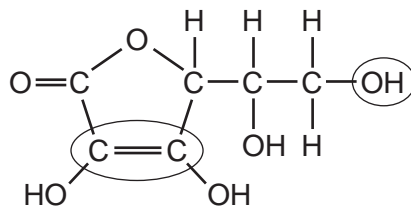
.....
..... [2]

(iii) Calculate its empirical formula and **then** its molecular formula.

The empirical formula of benzene is

The molecular formula of benzene is [2]

(c) The structural formula of Vitamin C is drawn below.



For
Examiner's
Use

(i) What is its molecular formula?

..... [1]

(ii) Name the two functional groups which are circled.

..... [2]

[Total: 12]

5 The electrolysis of concentrated aqueous sodium chloride produces three commercially important chemicals hydrogen, chlorine and sodium hydroxide.

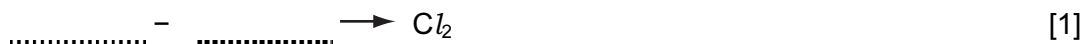
For
Examiner's
Use

(a) The ions present are $\text{Na}^+(\text{aq})$, $\text{H}^+(\text{aq})$, $\text{Cl}^-(\text{aq})$ and $\text{OH}^-(\text{aq})$.

(i) Complete the ionic equation for the reaction at the negative electrode (cathode).



(ii) Complete the ionic equation for the reaction at the positive electrode (anode).



(iii) Explain why the solution changes from sodium chloride to sodium hydroxide.

..... [1]

(b) (i) Why does the water supply industry use chlorine?

..... [1]

(ii) Name an important chemical that is made from hydrogen.

..... [1]

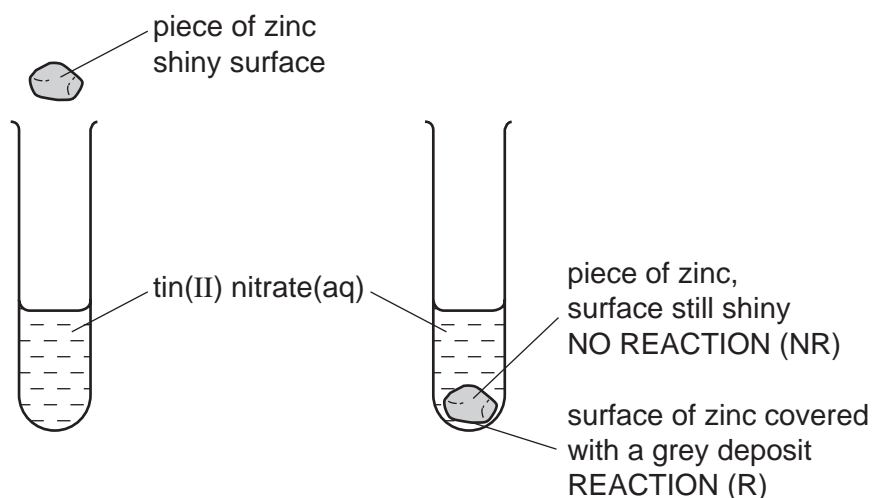
(iii) How is sodium hydroxide used to make soap?

..... [2]

[Total: 7]

6 The reactivity series lists metals in order of reactivity.

- (a) To find out which is the more reactive metal, zinc or tin, the following experiment could be carried out.



This experiment could be carried out with other metals and the results recorded in a table. Then the order of reactivity can be deduced.

- (i) The order was found to be:
- | | |
|-----------|----------------|
| manganese | most reactive |
| zinc | |
| tin | |
| silver | least reactive |

Complete the table of results from which this order was determined.

aqueous solution	tin Sn	manganese Mn	silver Ag	zinc Zn
tin(II) nitrate		R	NR	R
manganese(II) nitrate				
silver(I) nitrate				
zinc nitrate				

[3]

- (ii) Write the ionic equation for the reaction between tin atoms and silver(I) ions.

.....

[2]

(iii) The following is a redox reaction.



Indicate on the equation the change which is oxidation.
Give a reason for your choice.

..... [2]

(iv) Explain why experiments of this type cannot be used to find the position of aluminium in the reactivity series.

.....
..... [2]

(b) Potassium and calcium are very reactive metals at the top of the series. Because their ions have different charges, K^+ and Ca^{2+} , their compounds behave differently when heated.

(i) Explain why the ions have different charges.

.....
..... [2]

(ii) Their hydroxides are heated.
If the compound decomposes, complete the word equation.
If it does not decompose, write "no reaction".

Potassium hydroxide \longrightarrow

Calcium hydroxide \longrightarrow [2]

(iii) Complete the equations for the decomposition of their nitrates.

$2\text{KNO}_3 \longrightarrow$ +

$2\text{Ca}(\text{NO}_3)_2 \longrightarrow$ + + [4]

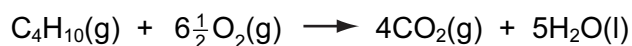
[Total: 17]

- 7 The alkanes are generally unreactive. Their reactions include combustion, substitution and cracking.

For
Examiner's
Use

(a) The complete combustion of an alkane gives carbon dioxide and water.

- (i) 10 cm³ of butane is mixed with 100 cm³ of oxygen, which is an excess. The mixture is ignited. What is the volume of unreacted oxygen left and what is the volume of carbon dioxide formed?



Volume of oxygen left = cm³

Volume of carbon dioxide formed = cm³ [2]

- (ii) Why is the incomplete combustion of any alkane dangerous, particularly in an enclosed space?

.....
..... [2]

(b) The equation for a substitution reaction of butane is given below.



- (i) Name the organic product.

..... [1]

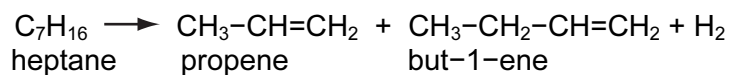
- (ii) This reaction does not need increased temperature or pressure. What is the essential reaction condition?

..... [1]

- (iii) Write a different equation for a substitution reaction between butane and chlorine.

..... [1]

- (c) Alkenes are more reactive and industrially more useful than alkanes. They are made by cracking alkanes.



- (i) Draw the structural formula of the polymer poly(propene).

[2]

- (ii) Give the structural formula and name of the alcohol formed when but-1-ene reacts with steam.

name

[1]

structural formula

[1]

- (iii) Deduce the structural formula of the product formed when propene reacts with hydrogen chloride.

[1]

[Total: 12]

DATA SHEET
The Periodic Table of the Elements

		Group											
I	II	III	IV	V	VI	VII	0						
		1 H Hydrogen 1							4 He Helium 2				
7 Li Lithium 3	9 Be Beryllium 4							20 Ne Neon 10					
23 Na Sodium 11	24 Mg Magnesium 12	27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulphur 16	35.5 Cl Chlorine 17	40 Ar Argon 18						
39 K Potassium 19	40 Ca Calcium 20	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36						
85 Rb Rubidium 37	88 Sr Strontium 38	101 Ru Ruthenium 44	106 Pd Palladium 46	112 Cd Cadmium 48	115 In Indium 49	127 I Iodine 53	131 Xe Xenon 54						
133 Cs Caesium 55	137 Ba Barium 56	190 Os Osmium 76	195 Pt Platinum 78	201 Hg Mercury 80	204 Tl Thallium 81	209 Bi Bismuth 83	210 Po Polonium 84						
226 Fr Francium 87	227 Ra Radium 88	232 Th Thorium 90	238 U Uranium 92	238 Np Neptunium 93	238 Pu Plutonium 94	238 Am Americium 95	238 Cm Curium 96						
*58-71 Lanthanoid series †90-103 Actinoid series													
140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	162 Dy Dysprosium 66	169 Tm Thulium 69						
175 Lu Lutetium 71	173 Yb Ytterbium 70	173 No Nobelium 102	167 Er Erbium 68	165 Ho Holmium 67	167 Fm Fermium 100	169 Md Mendelevium 101	175 Lr Lawrencium 103						

a	X	a = relative atomic mass X = atomic symbol b = proton (atomic) number
b	†	

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

University of Cambridge International Examinations is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CANDIDATE
NAME

CENTRE
NUMBER

--	--	--	--	--

CANDIDATE
NUMBER

--	--	--	--

* 7 5 0 7 3 1 6 8 3 0 *

CHEMISTRY

0620/32

Paper 3 (Extended)

October/November 2008

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES

Answer **all** questions.

A copy of the Periodic Table is printed on page 12.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part questions.

For Examiner's Use	
1	
2	
3	
4	
5	
6	
7	
Total	

This document consists of **12** printed pages.



1 Complete the following table.

gas	test for gas
	turns damp red litmus paper blue
	bleaches damp litmus paper
hydrogen	
oxygen	
carbon dioxide	

[Total: 5]

*For
Examiner's
Use*

2 There are three types of giant structure – ionic, metallic and macromolecular.

(a) Sodium sulphide is an ionic compound. Draw a diagram that shows the formula of the compound, the charges on the ions and the arrangement of the valency electrons around the negative ion.

Use x to represent an electron from a sodium atom.
Use o to represent an electron from a sulphur atom.

[3]

(b) (i) Describe metallic bonding.

.....
..... [3]

(ii) Use the above ideas to explain why
metals are good conductors of electricity,

..... [1]

metals are malleable.

..... [2]

(c) Silicon(IV) oxide has a macromolecular structure.

(i) **Describe** the structure of silicon(IV) oxide (a diagram is not acceptable).

.....
.....
..... [3]

(ii) Diamond has a similar structure and consequently similar properties.
Give **two** physical properties common to both diamond and silicon(IV) oxide.

.....
..... [2]

[Total: 14]

3 Steel is an alloy made from impure iron.

(a) Both iron and steel rust. The formula for rust is $\text{Fe}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$.
It is hydrated iron(III) oxide.

(i) Name the **two** substances that must be present for rusting to occur.

..... and [2]

(ii) Painting and coating with grease are two methods of preventing iron or steel from rusting. Give **two** other methods.

.....
..... [2]

(b) (i) Name a reagent that can reduce iron(III) oxide to iron.

..... [1]

(ii) Write a symbol equation for the reduction of iron(III) oxide, Fe_2O_3 , to iron.

..... [2]

(c) (i) Calculate the mass of one mole of $\text{Fe}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$.

..... [1]

(ii) Use your answer to (i) to calculate the percentage of water in rust.

.....
..... [2]

(d) Iron from the blast furnace is impure. Two of the impurities are carbon and silicon. These are removed by blowing oxygen through the molten iron and adding calcium oxide.

(i) Explain how the addition of oxygen removes carbon.

.....
..... [1]

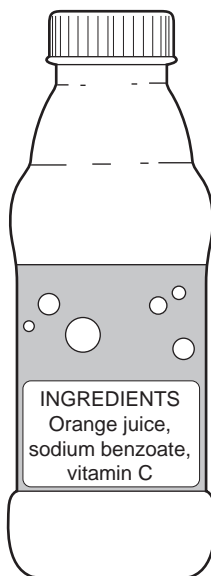
(ii) Explain how the addition of oxygen and calcium oxide removes silicon.

.....
..... [2]

[Total: 13]

- 4 Across the world, food safety agencies are investigating the presence of minute traces of the toxic hydrocarbon, benzene, in soft drinks. It is formed by the reduction of sodium benzoate by vitamin C.

For
Examiner's
Use



- (a) Sodium benzoate is a salt, it has the formula C_6H_5COONa . It can be made by the neutralisation of benzoic acid by sodium hydroxide.

- (i) Deduce the formula of benzoic acid.

..... [1]

- (ii) Write a word equation for the reaction between benzoic acid and sodium hydroxide.

..... [1]

- (iii) Name **two** other compounds that would react with benzoic acid to form sodium benzoate.

..... [2]

(b) Benzene contains 92.3% of carbon and its relative molecular mass is 78.

(i) What is the percentage of hydrogen in benzene?

..... [1]

(ii) Calculate the ratio of moles of C atoms: moles of H atoms in benzene.

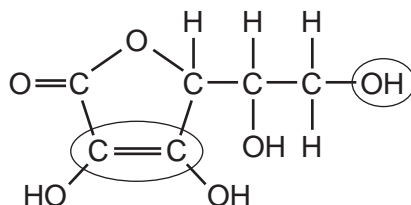
.....
..... [2]

(iii) Calculate its empirical formula and **then** its molecular formula.

The empirical formula of benzene is

The molecular formula of benzene is [2]

(c) The structural formula of Vitamin C is drawn below.



(i) What is its molecular formula?

..... [1]

(ii) Name the two functional groups which are circled.

..... [2]

[Total: 12]

- 5 The electrolysis of concentrated aqueous sodium chloride produces three commercially important chemicals; hydrogen, chlorine and sodium hydroxide.

For
Examiner's
Use

(a) The ions present are $\text{Na}^+(\text{aq})$, $\text{H}^+(\text{aq})$, $\text{Cl}^-(\text{aq})$ and $\text{OH}^-(\text{aq})$.

(i) Complete the ionic equation for the reaction at the negative electrode (cathode).



(ii) Complete the ionic equation for the reaction at the positive electrode (anode).



(iii) Explain why the solution changes from sodium chloride to sodium hydroxide.

..... [1]

(b) (i) Why does the water supply industry use chlorine?

..... [1]

(ii) Name an important chemical that is made from hydrogen.

..... [1]

(iii) Sodium hydroxide reacts with fats to make soap and glycerine
What type of compound are fats?

..... [1]

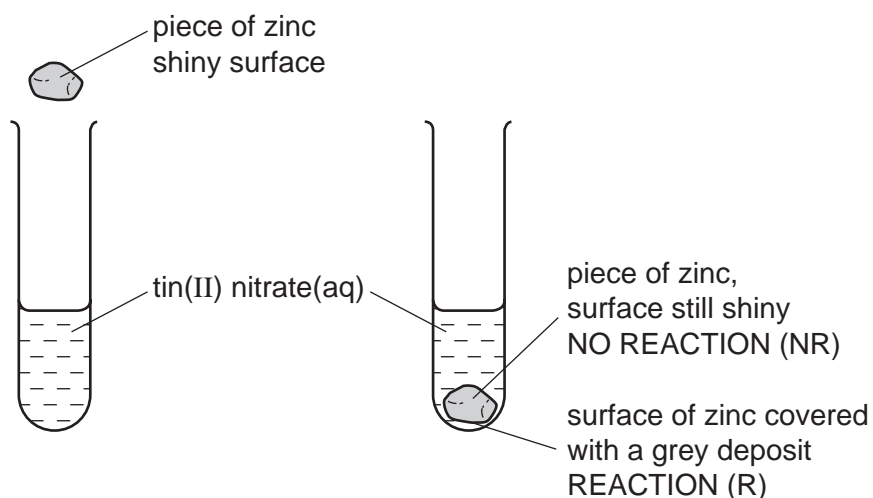
What type of the reaction is this?

..... [1]

[Total : 7]

6 The reactivity series lists metals in order of reactivity.

- (a) To find out which is the more reactive metal, zinc or tin, the following experiment could be carried out.



This experiment could be carried out with other metals and the results recorded in a table. Then the order of reactivity can be deduced.

- (i) The order was found to be:
 manganese most reactive
 zinc
 tin
 silver least reactive

Complete the table of results from which this order was determined.

aqueous solution	tin Sn	manganese Mn	silver Ag	zinc Zn
tin(II) nitrate		R	NR	R
manganese(II) nitrate				
silver(I) nitrate				
zinc nitrate				

[3]

- (ii) Write the equation for the reaction between zinc and silver(I) nitrate.

.....

[2]

- (iii) The following is a redox reaction.



Indicate on the equation which reagent is the oxidant or oxidizing agent.
Give a reason for your choice.

..... [2]

- (iv) Explain why experiments of this type cannot be used to find the position of aluminium in the reactivity series.

.....

..... [2]

- (b) Potassium and calcium are very reactive metals at the top of the series. Because their ions have different charges, K^+ and Ca^{2+} , their compounds behave differently when heated.

- (i) Explain why the ions have different charges.

.....

..... [2]

- (ii) Their hydroxides are heated.
If the compound decomposes, complete the word equation.
If it does not decompose, write "no reaction".

Potassium hydroxide \longrightarrow

Calcium hydroxide \longrightarrow [2]

- (iii) Complete the equations for the decomposition of their nitrates.

$2\text{KNO}_3 \longrightarrow$ +

$2\text{Ca}(\text{NO}_3)_2 \longrightarrow$ + + [4]

[Total: 17]

- 7 The alkanes are generally unreactive. Their reactions include combustion, substitution and cracking.

For
Examiner's
Use

(a) The complete combustion of an alkane gives carbon dioxide and water.

- (i) 20 cm³ of butane is mixed with 150 cm³ of oxygen, which is an excess. The mixture is ignited. What is the volume of unreacted oxygen left and what is the volume of carbon dioxide formed?



Volume of oxygen left = cm³

Volume of carbon dioxide formed = cm³ [2]

- (ii) Why is the incomplete combustion of any alkane dangerous, particularly in an enclosed space?

.....
..... [2]

(b) The equation for a substitution reaction of butane is given below.



- (i) Name the organic product.

..... [1]

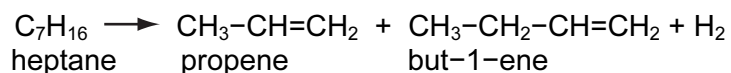
- (ii) This reaction does not need increased temperature or pressure. What is the essential reaction condition?

..... [1]

- (iii) Write a different equation for a substitution reaction between butane and chlorine.

..... [1]

- (c) Alkenes are more reactive and industrially more useful than alkanes. They are made by cracking alkanes.



- (i) Draw the structural formula of the polymer poly(propene).

[2]

- (ii) Give the structural formula and name of the alcohol formed when propene reacts with steam.

name

[1]

structural formula

[1]

- (iii) Deduce the structural formula of the product formed when but-1-ene reacts with hydrogen chloride.

[1]

[Total: 12]

For
Examiner's
Use

DATA SHEET
The Periodic Table of the Elements

		Group										
I	II	III	IV	V	VI	VII	0					
		1 H Hydrogen 1						4 He Helium 2				
7 Li Lithium 3	9 Be Beryllium 4							20 Ne Neon 10				
23 Na Sodium 11	24 Mg Magnesium 12	5 B Boron 5	12 C Carbon 6	14 N Nitrogen 7	16 O Oxygen 8	19 F Fluorine 9		35.5 Cl Chlorine 17				
39 K Potassium 19	40 Ca Calcium 20	13 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulphur 16	35.5 Cl Chlorine 17	40 Ar Argon 18					
85 Rb Rubidium 37	88 Sr Strontium 38	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36					
133 Cs Caesium 55	137 Ba Barium 56	65 Zn Zinc 30	112 Cd Cadmium 48	115 In Indium 49	122 Sb Antimony 51	127 I Iodine 53	131 Xe Xenon 54					
226 Ra Radium 88	227 Ac Actinium 89	204 Pb Lead 82	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 Po Polonium 84	210 Po Polonium 84	210 Po Polonium 84				
		59 Ni Nickel 28	64 Cu Copper 29	106 Pd Palladium 46	195 Pt Platinum 78	197 Au Gold 79						
		56 Fe Iron 26	59 Co Cobalt 27	108 Ag Silver 47	201 Hg Mercury 80							
		55 Mn Manganese 25	56 Fe Iron 26	112 Cd Cadmium 48								
		52 Cr Chromium 24	56 Fe Iron 26	112 Cd Cadmium 48								
		51 V Vanadium 23	56 Fe Iron 26	112 Cd Cadmium 48								
		48 Ti Titanium 22	56 Fe Iron 26	112 Cd Cadmium 48								
		45 Sc Scandium 21	56 Fe Iron 26	112 Cd Cadmium 48								
		89 Y Yttrium 39	56 Fe Iron 26	112 Cd Cadmium 48								
		139 La Lanthanum 57	56 Fe Iron 26	112 Cd Cadmium 48								
		178 Hf Hafnium 72	56 Fe Iron 26	112 Cd Cadmium 48								
		181 Ta Tantalum 73	56 Fe Iron 26	112 Cd Cadmium 48								
		184 W Tungsten 74	56 Fe Iron 26	112 Cd Cadmium 48								
		186 Re Rhenium 75	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		144 Nd Neodymium 60	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59	56 Fe Iron 26	112 Cd Cadmium 48								
		141 Pr Praseodymium 59										