
CHEMISTRY**0620/43**

Paper 4 Extended Theory

October/November 2018

MARK SCHEME

Maximum Mark: 80

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the October/November 2018 series for most Cambridge IGCSE™, Cambridge International A and AS Level components and some Cambridge O Level components.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Question	Answer	Marks
1(a)	oxygen	1
1(b)	hematite	1
1(c)	sulfur dioxide	1
1(d)	ammonia	1
1(e)	carbon monoxide	1
1(f)	sodium chloride	1
1(g)	carbon dioxide	1
1(h)	oxygen	1

Question	Answer	Marks
2(a)(i)	M1 breakdown of an ionic compound when molten or in aqueous solution M2 (using) electricity / electric current	2
2(a)(ii)	M1 electron(s) M2 ion(s)	2
2(b)(i)	M1 inert / unreactive M2 conducts electricity	2

Question	Answer				Marks
2(b)(ii)	observation at anode(+)	name of product at anode(+)	observation at cathode(-)	name of product at cathode(-)	6
	M1 green / yellow bubbles	M2 chlorine		M3 hydrogen	
		M4 oxygen	M5 pink / brown solid	M6 copper	

Question	Answer	Marks
3(a)	$[(64 \times 2) + 56 + 119 + (32 \times 4) =] \mathbf{431}$	1
3(b)	$[(119 / 151) \times 100 =] \mathbf{78.8 (\%)}$	1
3(c)	SnO ₂ because the percentage of tin is larger in SnO ₂ or answer to (b) > 27.6 %	1
3(d)	SnO ₂ + 2C → Sn + 2CO M1 all formulae correct M2 equation fully correct	2
3(e)	M1 (→) Fe ²⁺ + Sn OR 2Fe + 3Sn ²⁺ → 2Fe ³⁺ + 3Sn M2 (→) Sn ²⁺ + Cu OR Sn + 2Cu ²⁺ → Sn ⁴⁺ + 2Cu	2
3(f)(i)	M1 glowing splint M2 relights / rekindles	2

Question	Answer	Marks						
3(f)(ii)	M1 nitrogen dioxide / nitrogen(IV) oxide M2 brown (gas)	2						
3(f)(iii)	$2\text{Cu}(\text{NO}_3)_2 \rightarrow 2\text{CuO} + 4\text{NO}_2 + \text{O}_2$	1						
3(g)(i)	zinc acts as a barrier which prevents contact between iron and water and air / oxygen	1						
3(g)(ii)	<p>SUMMARY</p> <table border="1" style="margin-left: auto; margin-right: auto;"> <tbody> <tr> <td>M1</td> <td>comparison of reactivity</td> </tr> <tr> <td>M2</td> <td>zinc loses electrons</td> </tr> <tr> <td>M3</td> <td>where electrons move to OR iron does not lose electrons</td> </tr> </tbody> </table> <p>M1 zinc is more reactive than iron / steel ORA</p> <p>M2 zinc loses electrons / zinc is oxidised</p> <p>M3 electrons are transferred to iron / iron is not oxidised / iron does not lose electrons</p>	M1	comparison of reactivity	M2	zinc loses electrons	M3	where electrons move to OR iron does not lose electrons	3
M1	comparison of reactivity							
M2	zinc loses electrons							
M3	where electrons move to OR iron does not lose electrons							

Question	Answer	Marks
4(a)	M1 (Mol KOH =) $0.00125 / 1.25 \times 10^{-3}$ M2 (Mol H ₂ SO ₄ =) $0.000625 / 6.25 \times 10^{-4}$ M3 (Conc H ₂ SO ₄ =) $0.03125 / 3.125 \times 10^{-2}$ (mol / dm ³)	3

Question	Answer	Marks										
4(b)	<p>SUMMARY</p> <table border="1" data-bbox="631 284 1641 611"> <tr> <td>M1</td> <td>repeat</td> </tr> <tr> <td>M2</td> <td>heat (liquid or solution should be implied)</td> </tr> <tr> <td>M3</td> <td>when to stop heating</td> </tr> <tr> <td>M4</td> <td>what to do after heating</td> </tr> <tr> <td>M5</td> <td>method of drying crystals (crystals or solid should be implied)</td> </tr> </table> <p>M1 repeat without indicator using same volumes</p> <p>M2 evaporate / heat / warm / boil / leave in sun</p> <p>M3 until most of the water is gone / some water left / saturation(point) / crystallisation point / evaporate some of the water</p> <p>M4 leave / (allow to) cool / allow to crystallise</p> <p>M5 details of drying</p>	M1	repeat	M2	heat (liquid or solution should be implied)	M3	when to stop heating	M4	what to do after heating	M5	method of drying crystals (crystals or solid should be implied)	5
M1	repeat											
M2	heat (liquid or solution should be implied)											
M3	when to stop heating											
M4	what to do after heating											
M5	method of drying crystals (crystals or solid should be implied)											
4(c)(i)	<p>M1 bubbles / effervescence / fizzing</p> <p>M2 solid or magnesium dissolves / solid or magnesium disappears</p>	2										
4(c)(ii)	lilac flame	1										
4(c)(iii)	white precipitate	1										
4(d)(i)	<p>$\text{Mg(OH)}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + 2\text{H}_2\text{O}$</p> <p>M1 formula of both Mg(OH)_2 and MgSO_4</p> <p>M2 equation fully correct</p>	2										

Question	Answer	Marks
4(d)(ii)	$\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$ <p>M1 formula of ZnSO_4</p> <p>M2 equation fully correct</p>	2
4(d)(iii)	$\text{Na}_2\text{CO}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{CO}_2 + \text{H}_2\text{O}$ <p>M1 formulae of both Na_2CO_3 and Na_2SO_4</p> <p>M2 equation fully correct</p>	2

Question	Answer	Marks
5(a)	<p>M1 volume of gas</p> <p>M2 time</p>	2
5(b)	<p>M1 rate decreases / reaction gets slower</p> <p>M2 concentration of acid decreases</p> <p>M3 fewer collisions per unit time</p>	3

Question	Answer	Marks
5(c)	<p>M1 particles have more kinetic energy</p> <p>M2 particles move faster</p> <p>M3 more collisions per unit time</p> <p>M4 more of the particles have energy greater than or equal to activation energy / more of the collisions have energy greater than or equal to activation energy</p> <p>OR</p> <p>more of the particles have sufficient energy to react / more of the collisions have sufficient energy to react</p> <p>OR</p> <p>A greater percentage or greater proportion or greater fraction of collisions are successful</p>	4
5(d)	<p>ANY TWO FROM:</p> <ul style="list-style-type: none"> • increase concentration of hydrochloric acid • decrease particle size of calcium carbonate / increase surface area of calcium carbonate • (add)catalyst 	2

Question	Answer	Marks						
6(a)(i)	<p>SUMMARY</p> <table border="1" data-bbox="927 284 1348 480"> <tr> <td>M1 and M4</td> <td>reactants</td> </tr> <tr> <td>M2 and M5</td> <td>conditions</td> </tr> <tr> <td>M3 and M6</td> <td>equation</td> </tr> </table> <p>FERMENTATION: M1 glucose / sucrose / starch / other named carbohydrate can score in equation as correct formula M2 Zymase / Yeast / 37°C M3 $C_6H_{12}O_6 \rightarrow 2C_2H_5OH + 2CO_2$</p> <p>HYDRATION: M4 Ethene and steam or water can score in equation as correct formulae M5 H_3PO_4 (catalyst) / 300°C / 60 atm M6 $C_2H_4 + H_2O \rightarrow C_2H_5OH$</p>	M1 and M4	reactants	M2 and M5	conditions	M3 and M6	equation	6
M1 and M4	reactants							
M2 and M5	conditions							
M3 and M6	equation							
6(a)(ii)	<p>ANY TWO FROM:-</p> <ul style="list-style-type: none"> • carbohydrates are renewable • fossil fuels are non-renewable • lower temperature means fossil fuels conserved ORA • lower temperature means lower energy costs ORA • hydration reaches an equilibrium meaning lower yield ORA 	2						
6(a)(iii)	<p>M1 solvent</p> <p>M2 fuel</p>	2						
6(b)(i)	E	1						

Question	Answer	Marks
6(b)(ii)	D	1
6(b)(iii)	B	1
6(b)(iv)	C	1
6(b)(v)	A	1