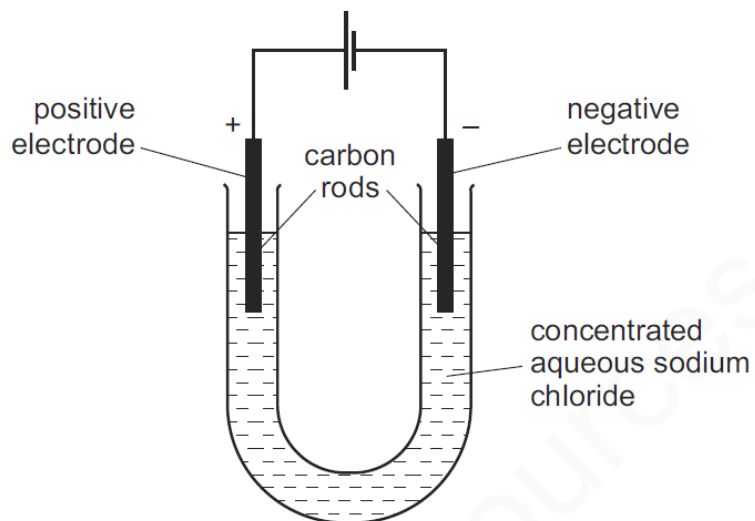


NO:

ELECTROLYSIS-CONC NaCl-SET-1

1

Electricity is passed through concentrated aqueous sodium chloride, as shown.

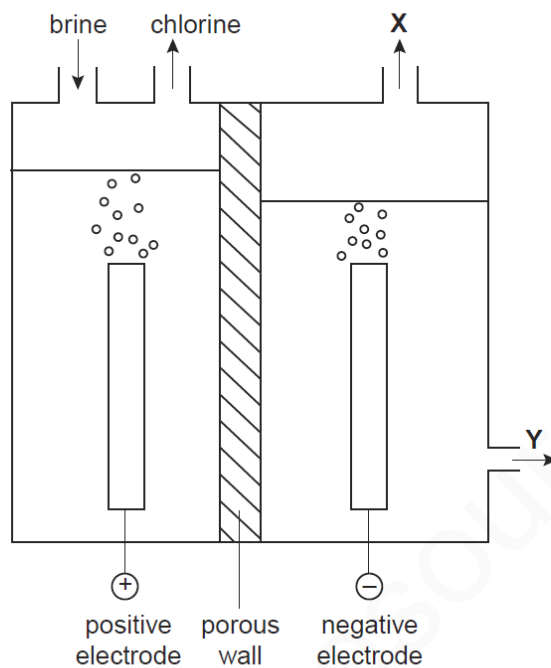


What is the test for the gas formed at the positive electrode?

- A bleaches damp litmus paper
- B 'pops' with a lighted splint
- C relights a glowing splint
- D turns damp red litmus paper blue

2

The diagram represents the electrolysis of brine (aqueous sodium chloride).



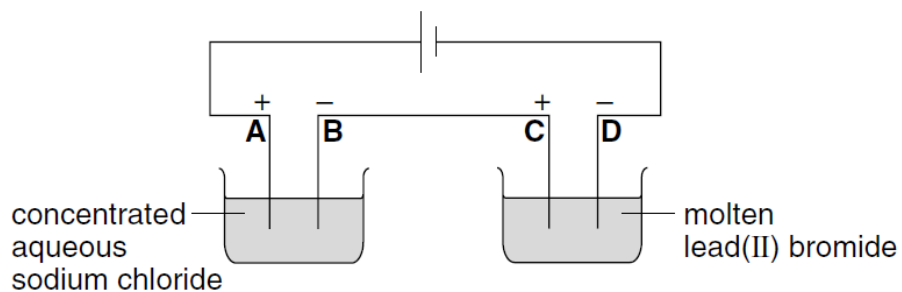
What are products **X** and **Y**?

	X	Y
A	hydrogen	aqueous sodium hydroxide
B	hydrogen	hydrochloric acid
C	oxygen	aqueous sodium hydroxide
D	oxygen	hydrochloric acid

3

The following electrolysis circuit is set up, using inert electrodes.

At which electrode is a metal deposited?



4

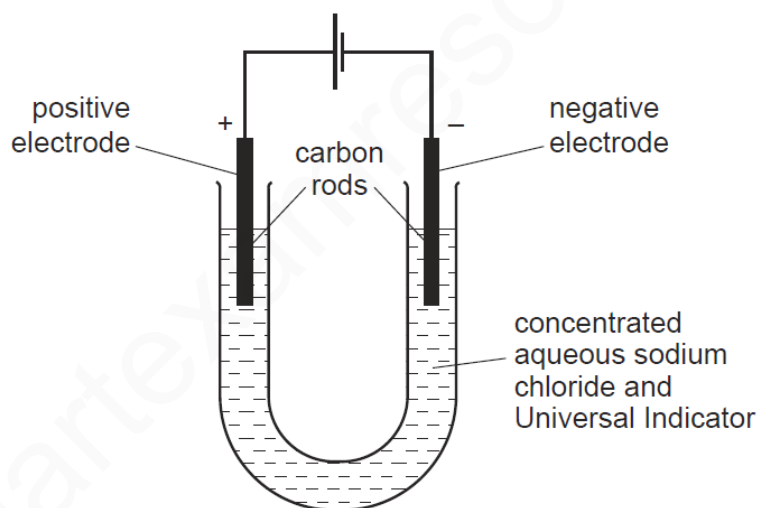
The electrolysis of concentrated aqueous sodium chloride makes three products.

Which products are shown at the correct electrodes?

	anode (+ve)	cathode (-ve)
A	chlorine	sodium hydroxide
B	sodium hydroxide	chlorine
C	hydrogen	sodium
D	sodium	hydrogen

5

The diagram shows the electrolysis of concentrated aqueous sodium chloride.

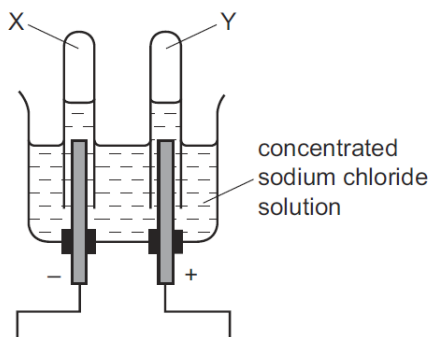


What is the colour of the Universal Indicator at each electrode after five minutes?

	colour at anode (+ electrode)	colour at cathode (- electrode)
A	blue/purple	red
B	red	blue/purple
C	red	colourless
D	colourless	blue/purple

6

When concentrated sodium chloride solution is electrolysed, elements X and Y are formed.

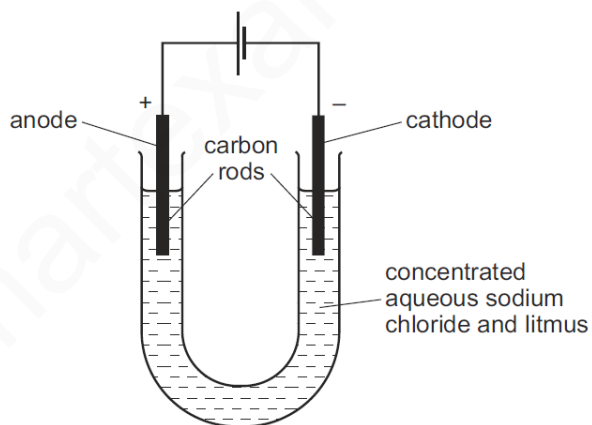


What are X and Y?

	X	Y
A	chlorine	hydrogen
B	hydrogen	chlorine
C	hydrogen	oxygen
D	oxygen	hydrogen

7

The diagram shows the electrolysis of concentrated aqueous sodium chloride.



What is the colour of the litmus at each electrode after five minutes?

	colour at anode	colour at cathode
A	blue	red
B	red	blue
C	red	colourless
D	colourless	blue